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When an ionic substance dissolves in water, the positive and negative ions separate and become hydrated (they interact with water molecules rather than each other).

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More Exam Questions on 1.2 Amount of Substance (mark scheme) 1.2 Exercise 1 - using moles and reacting masses 1.2 Exercise 2 - solutions 1.2 Exercise 3 - ideal gas equation 1.2 Exercise 4 - empirical and molecular formulae 1.2 Exercise 5 - ionic equations 1.2 Exercise 6 - more complex calculations

1.2 Amount Of Substance - A-Level Chemistry

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AS level Chemistry - Amount of substance | Teaching Resources

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Welcome to Topic 2 - AMOUNT OF SUBSTANCE. Topic 2 specification content Topic 2 notes

Topic 2 - Amount of Substance - A-Level Chemistry

• The amount of energy needed to make 1 g of a substance 1°C (1 K) hotter is called the specific heat capacity (measured in J g⁻¹ K⁻¹). • The following equation is then used to find the amount of heat energy give out (or absorbed). $q = m c \Delta T$ $q =$ heat energy given out (J) $m =$ mass of substance heated (g) $\Delta T =$ temperature rise (K)

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Worksheets and lesson ideas to challenge students to think about mole calculations, titrations and equations (GCSE and Key Stage 3) Students must master the mole and other quantitative aspects if they are to excel in chemistry and fall in love with ... Amounts of substance teaching resources Read More »

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Chemsheets AS 008 (Amount of substance) READ. 1) A compound contains 59.4% carbon, 10.9% hydrogen, 13.9% nitrogen and 15.8% oxygen, by

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mass. Find the empirical formula of the compound. 2) A compound containing carbon, hydrogen and oxygen only contains 74.2% carbon and 7.9% hydrogen. Its Mr is . found to be 178 by mass spectroscopy. ...

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The amount of substance is sometimes referred to as the chemical amount. The mole (symbol: mol) is a unit of amount of substance in the International System of Units, defined (since 2019) by fixing the Avogadro constant at the given value. Historically, the mole was defined as the amount of substance in 12 grams of the carbon-12 isotope.

Amount of substance - Wikipedia

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that only need a small amount of energy to overcome c Explain why pure water does not conduct electricity. no charged particles that can move so cannot carry charge through the substance d Explain what the molecular formula H₂O means. 2 H atoms and 1 O atom in each molecule 4 Sodium oxide is an ionic substance with the formula Na₂O.

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